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- While you should practice working as fast as possible, it is more important at this point in the course, that you practice without a calculator, even if it slows you down. Look for the “easy math” – common factors and rough estimation – do not do “long division” to try to get exact values. Remember it is a MC test, use the answers

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Stoichiometry - Practice Test Questions & Chapter Exam ...

Lots and lots and lots of practice problems with mole ratios. This is the first step in learning stoichiometry, for using a chemical equation to get mole ratios and using conversion factors and ...

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352 Chapter 12 Visit the Chemistry Web site at chemistrymc.com to find links to stoichiometry. 12.1 What is stoichiometry? 353 DISCOVERY LAB Materials 10-mL graduated cylinder 100-mL beaker stirring rod 0.01M potassium perman-ganate (KMnO₄) 0.01M sodium hydrogen sulfite (NaHSO₃) Observing a Chemical Reaction Reactants are consumed in a chemical reaction as products are produced. What ...

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Mole Ratio Practice Problems

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Chapter 12: Stoichiometry - Laramie County School District #2

Chapter 12 . Stoichiometry Notes Packet Big Picture Ideas: The identity of the reactants helps scientists to predict the products in a chemical reaction. Quantitative relationships exist with all chemical reactions that allow scientists to predict amounts of products formed, reactants consumed, and percent yield based on theoretical maximum.

CHAPTER 11: STOICHIOMETRY

Chapter 1 - Intro to Chemistry. Chapter 2- Classification of Matter. Chapter 3 - Math and Density.

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Chapter 4 - Atomic Theory. Chapter 6 - The Periodic Table and Trends . Chapter 9 - Chemical Names and Formulas. Chapter 10 - Chemical Quantities. Chapter 11a - Writing, Balancing, and Reaction Types. Chapter 11b - Predicting Chemical Equations. Chapter 25 - Nuclear Chemistry. Chapter 20 ...