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CHAPTER 6 - ELECTRONIC STRUCTURE AND THE PERIODIC TABLE

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- The Photoelectric
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replaced AS 90522.
Prior to 2013, this was
an external standard -

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AS90522 Atoms, Photons and Nuclei. It is likely to be assessed using an internal test from 2013 onwards (although teachers can select from a range of assessment techniques). There were only minor ...

LIGHT QUESTIONS and ANSWERS - Santa Monica College

(ground state) to
higher energy levels

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(excited states). Energy is released as electrons return to their lower energy levels. 18. Is energy absorbed or released for the electron transition shown in the diagram to the right? Explain.
absorb* Electron Energy and Light 5432

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Answers and
Berühmt ...**

LIGHT QUESTIONS and
Page 6/24

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ANSWERS 1. What does it mean to say an atom is excited? An electron or electrons have more energy than their ground state, which is where they would be if they had not gained energy in some way. 2. How does the energy of the photon that is emitted by an electron jumping between two

**Level 3 Physics:
Atoms - The**

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**Photoelectric Effect
- Answers**

Answer and

Explanation: a) The energy of the photon and the electron are equal,
$$\{eq\}E_{\{ph\}}=E_{\{e\}}=E_{\{/eq\}}.$$
 Their associated wavelengths are also equal,

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Answers Along with
...**

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**Suppose an electron
and a photon have
the same energy
and ...**

An electron in a
hydrogen atom relaxes
to the $n=4$ level,
emitting light of 114

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Key

THz. What is the value of n for the level in which the electron originated? Express answer as integer.

6.2: Quantized Energy and Photons - Chemistry LibreTexts

AP CHEMISTRY
CHAPTER REVIEW
CHAPTER 6:
ELECTRONIC
STRUCTURE AND THE
PERIODIC TABLE You
should be familiar with

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the wavelike properties of light: frequency (ν), wavelength (λ), and energy (E) as well as the equations that show their relationships ($E = h\nu$ and $c = \lambda\nu$) You should be familiar with the electromagnetic spectrum and atomic spectra (bright-line spectra).

Electronvolt - Wikipedia

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If photons of light with

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energy less than E_0 strike a metal surface, no single photon has enough energy to eject an electron, so no electrons are emitted regardless of the intensity of the light. If a photon with energy greater than E_0 strikes the metal, then part of its energy is used to overcome the forces that hold the electron to the ...

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In this section we will discuss the energy level of the electron of a hydrogen atom, and how it changes as the electron undergoes transition. According to Bohr's theory, electrons of an atom revolve around the nucleus on certain orbits, or electron shells. Each orbit has its specific energy

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level, which is expressed as a negative value. This is because the electrons on the orbit are "captured ...

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Page 19/24

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**Energy Level and
Transition of
Electrons | Brilliant
Math ...**

In physics, an electronvolt (symbol eV, also written electron-volt and electron volt) is the amount of kinetic energy gained (or lost) by a single electron

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accelerating from rest through an electric potential difference of one volt in vacuum. When used as a unit of energy, the numerical value of 1 eV in joules (symbol J) is equivalent to the numerical value of the charge of an electron in ...

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Key
the hydrogen spectrum
in Model 2. a. Which
color of light has the
most energy? b. Which
color of light has the
least energy? 12. Does
a gas discharge tube
filled with boron emit
the same wavelengths
of light as a tube filled
with hydrogen? Use
evidence from Model 2
to support your
answer. 13.

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Page 22/24

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Quizizz

Electron Energy and

Light 3 11. Consider the hydrogen spectrum in Model 2. a. Which

color of light has the most energy? b. Which

color of light has the least energy? 12. Does

a gas discharge tube filled with boron emit

the same wavelengths of light as a tube filled

with hydrogen? Use evidence from Model 2

to support your answer. 13.

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For an electron to move from an energy level close to the nucleus to an energy level far from the nucleus it would need to ____ energy. (#16)
answer choices