

## Modern Chemistry Chapter 3 Section Review Answers

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CHAPTER 10 REVIEW States of Matter SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Match description on the right to the correct crystal type on the left. ... SECTION 5 MODERN CHEMISTRY STATES OF MATTER. STATES OF MATTER MODERN CHEMISTRY.

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see figure 3-3 page 67 . C: Modern Atomic Theory. Some of Dalton's postulates have been proven to be incorrect. His third postulate indicated that atoms cannot be subdivided. His second postulate indicates atoms of the same element have identical mass. Isotopes. His theory has been modified -- not discarded. Homework: Section 3-1

### 3 Atoms: The Building Blocks of Matter

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### 10 States of Matter - Ms. Agostine's Chemistry Page

MODERN CHEMISTRY The atomic number of nickel-60 is 28. How many neutrons does this iso- tope have? Carbon-14 has 8 neutrons. What is the atomic number of carbon-14? 11 CHAPTER 3 TEST HR W material copyrighted under notice appealing earlier in this work.

### 1 Matter and Change - HUBBARD'S HONORS CHEMISTRY CLASS

Chapter 6: Chemical Bonding, Modern Chemistry. This is a working presentation of the notes for this chapter. ... Chemical bonding chapter 6 1. Chapter 6 - ModernCHEMICAL BONDING Chemistry S.Martinez ... IONIC BONDS AND Chapter 6 Section 3 COMPOUNDS11/19/2010 S.Martinez 39 40. IONIC BONDING & IONIC COMPOUNDS Most of the rocks & minerals in the ...

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### 5 The Periodic Law

3:1 mol N<sub>2</sub>; 2 mol NH<sub>3</sub>; 3 mol H<sub>2</sub>; or their reciprocals 7. Given the following equation: 4NH<sub>3</sub>(g) + 6NO(g) → 5N<sub>2</sub>(g) + 6H<sub>2</sub>O(g) 1 mol NO:1 mol H<sub>2</sub>O a. What is the mole ratio of NO to H<sub>2</sub>O? 3 mol NO:2 mol NH<sub>3</sub> b. What is the mole ratio of NO to NH<sub>3</sub>? 0.360 mol c. If 0.240 mol of NH<sub>3</sub> react according to the above equation, how many moles of NO will ...

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CHAPTER 1 REVIEW Matter and Change MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Classify each of the following as a homogeneous or heterogeneous substance. homogeneous a. sugar homogeneous d. plastic wrap homogeneous b. iron filings heterogeneous e. cement sidewalk heterogeneous c. granola bar 2. For each type of investigation, select the most appropriate ...

### New Page 1 [srvhs.org]

CHAPTER 6 REVIEW Chemical Bonding SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. a The notation for sodium chloride, NaCl, stands for one (a) formula unit. (c) crystal. (b) molecule. (d) atom. 2. d In a crystal of an ionic compound, each cation is surrounded by a number of (a) molecules. (c) dipoles. (b) ...

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### Modern Chemistry Chapter 6 Section 3 Flashcards | Quizlet

CHAPTER 3 REVIEW Atoms: The Building Blocks of Matter SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Explain the difference between the mass number and the atomic number of a nuclide. Mass number is the total number of protons and neutrons in the nucleus of an isotope.

### Holt's Modern Chemistry Textbook | CourseNotes

CHAPTER 3 REVIEW Atoms: The Building Blocks of Matter SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Explain the difference between the mass number and the atomic number of a nuclide. Mass number is the total number of protons and neutrons in the nucleus of an isotope.

### Modern Chemistry (Florida 2015-2016) Chapter 3 Section 3 ...

if two or more different compounds are composed of the same two elements, then the ratio of the masses of the second element combined with a certain mass of the first element is always a ratio of small whole numbers; textbook: if two or more different compounds are composed of the same two elements with a certain mass of the first element is always a ratio of small whole numbers

### 6 Chemical Bonding

CHAPTER 5 REVIEW The Periodic Law SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. c In the modern periodic table, elements are ordered (a) according to decreasing atomic mass. (b) according to Mendeleev's original design. (c) according to increasing atomic number. (d) based on when they were discovered. 2. d Mendeleev noticed that certain similarities in the ...

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