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### **Problem Solutions Of Chemical Thermodynamic**

Problem : Given that the free energy of formation of liquid water is  $-237 \text{ kJ / mol}$ , calculate the potential for the formation of hydrogen and oxygen from water. To solve this problem we must first calculate  $\Delta G$  for the reaction, which is  $-2 (-237 \text{ kJ / mol}) = 474 \text{ kJ / mol}$ . Knowing that  $\Delta G = -nFE$  and  $n = 4$ , we calculate the potential is  $-1.23 \text{ V}$ .

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chapter 06: thermodynamic relations. chapter 07: ideal and real gas processes and relations. chapter 08: vapor power and refrigeration cycles. chapter 09: air-standard power and refrigeration cycles. chapter 10: mixtures and solutions. chapter 11: chemical reactions and equilibrium

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**The first law of thermodynamics - problems and solutions ...**

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**Problems in Chemical Thermodynamics, With Solutions**

The first law of thermodynamics - problems and solutions. 1. 3000 J of heat is added to a system and 2500 J of work is done by the system. What is the change in internal energy of the system? Known : Heat (Q) = +3000 Joule. Work (W) = +2500 Joule . Wanted: the change in internal energy of the system Solution :

